Regional Symposium on Electrochemistry
RSE-SEE
South-East Europe
Graz, Austria, 11–15 July 2022

Penetration of Cl^- ions from solution into an organic self-assembled monolayer on a metal substrate: Trends and modeling aspects

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Chemie ParisTech

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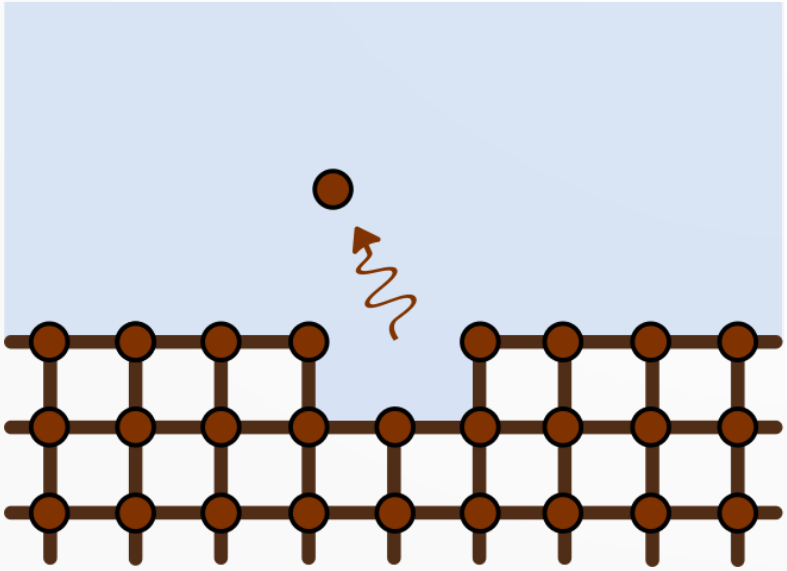
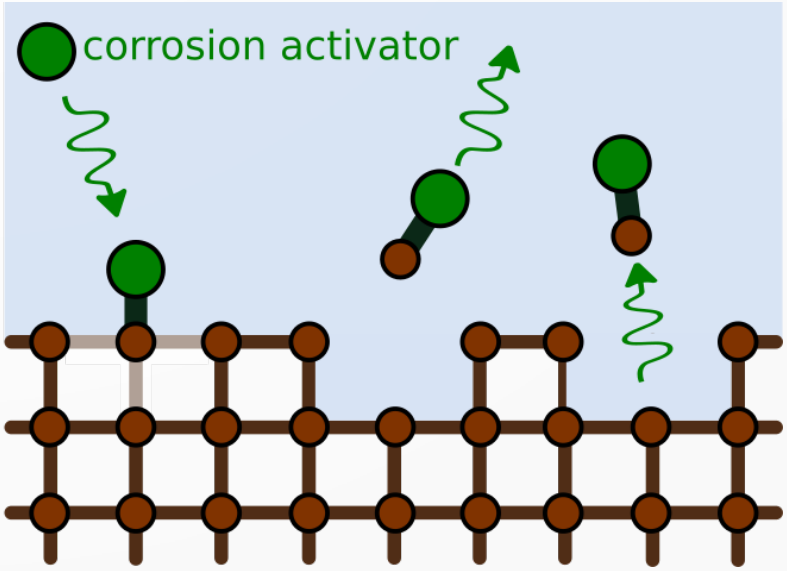
Introducing the context

Corrosion of metals: $M \rightarrow M^{z+} + ze^{-}$

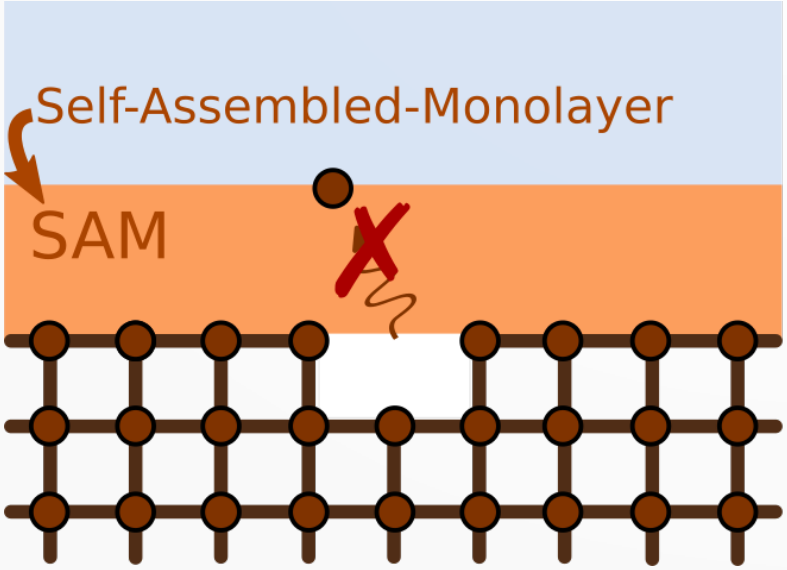


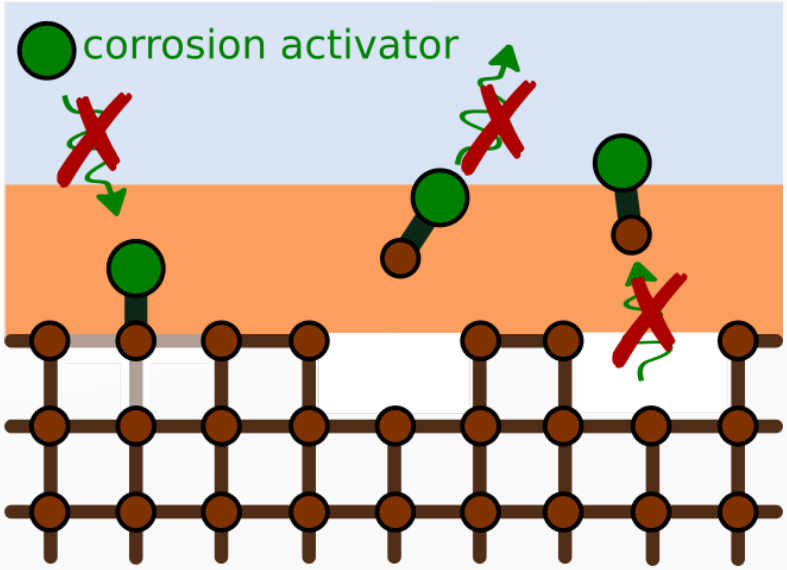


Introducing the context

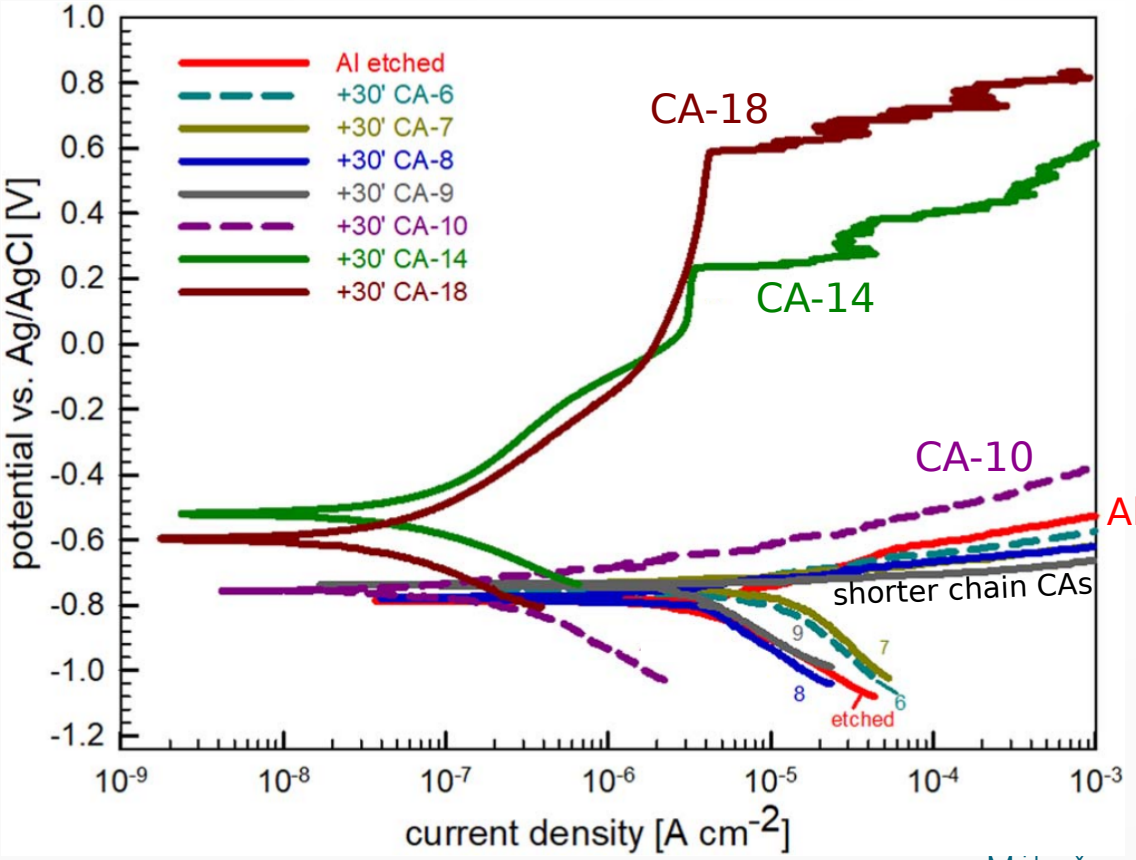
Corrosion of metals: $M \rightarrow M^{z+} + ze^-$

corrosion can be reduced with surfactant inhibitors that form a self-assembled-monolayer (SAM) on the surface of a metal





Carboxylic acids (CA) as inhibitors of Al



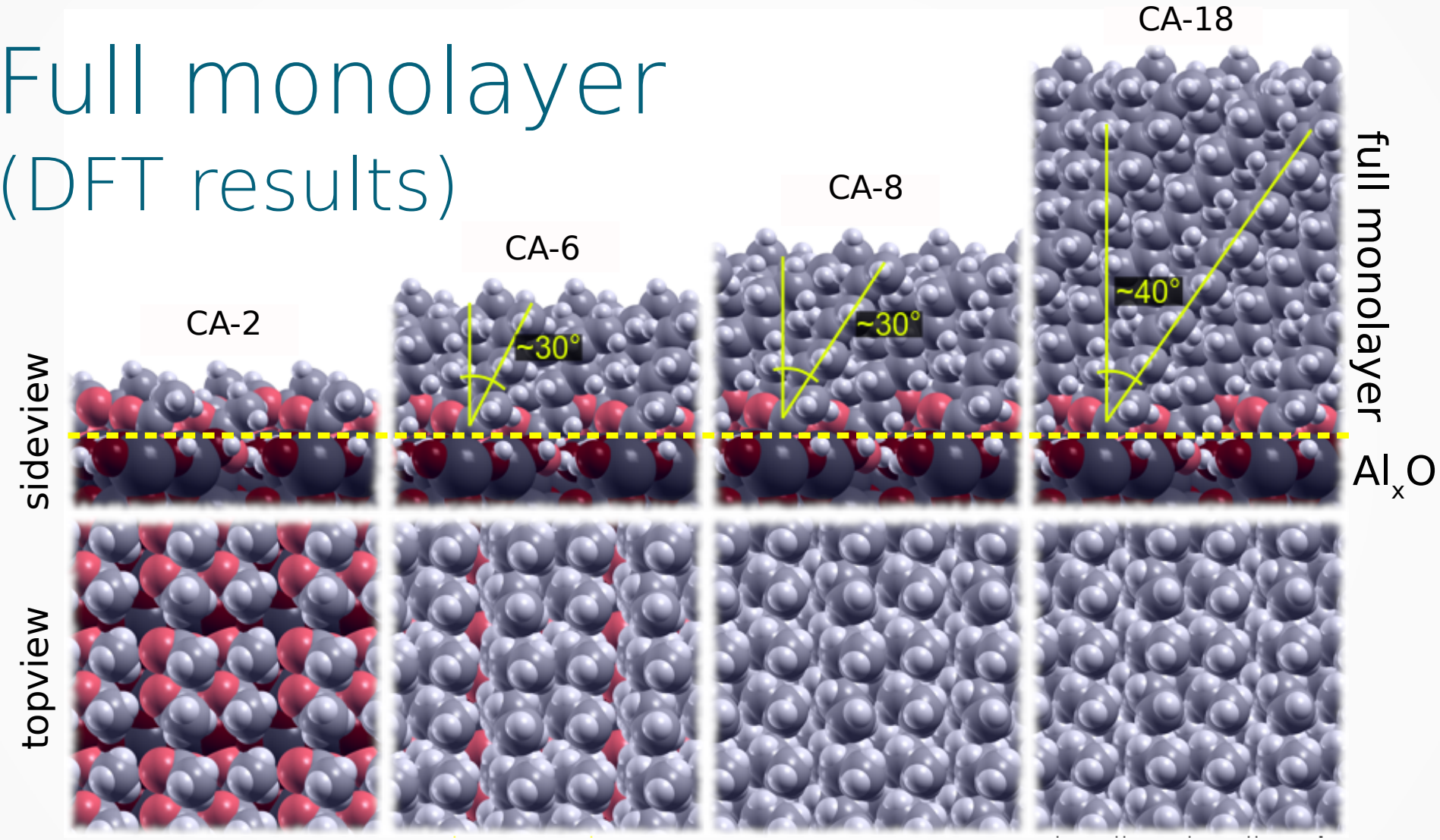
CA-n = $C_nH_{2n}O_2$
(CA-8 = octanoic acid)

only long-chain CAs are efficient inhibitors

WHY?

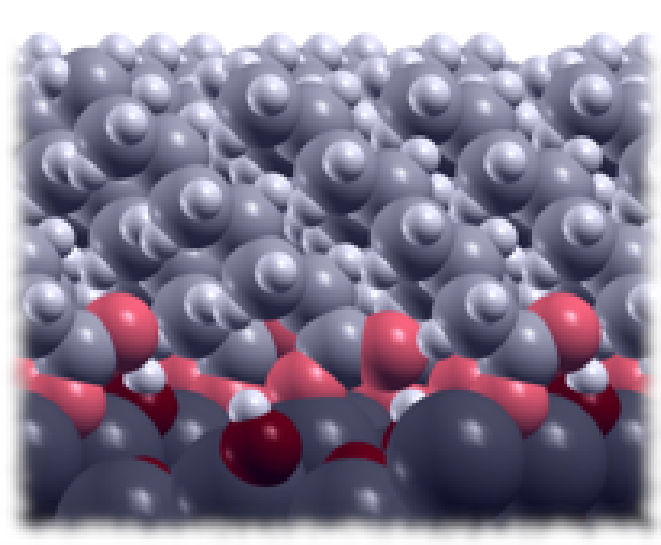
Milošev et al., J. Electrochem. Soc. **166**, C3131

Full monolayer
(DFT results)

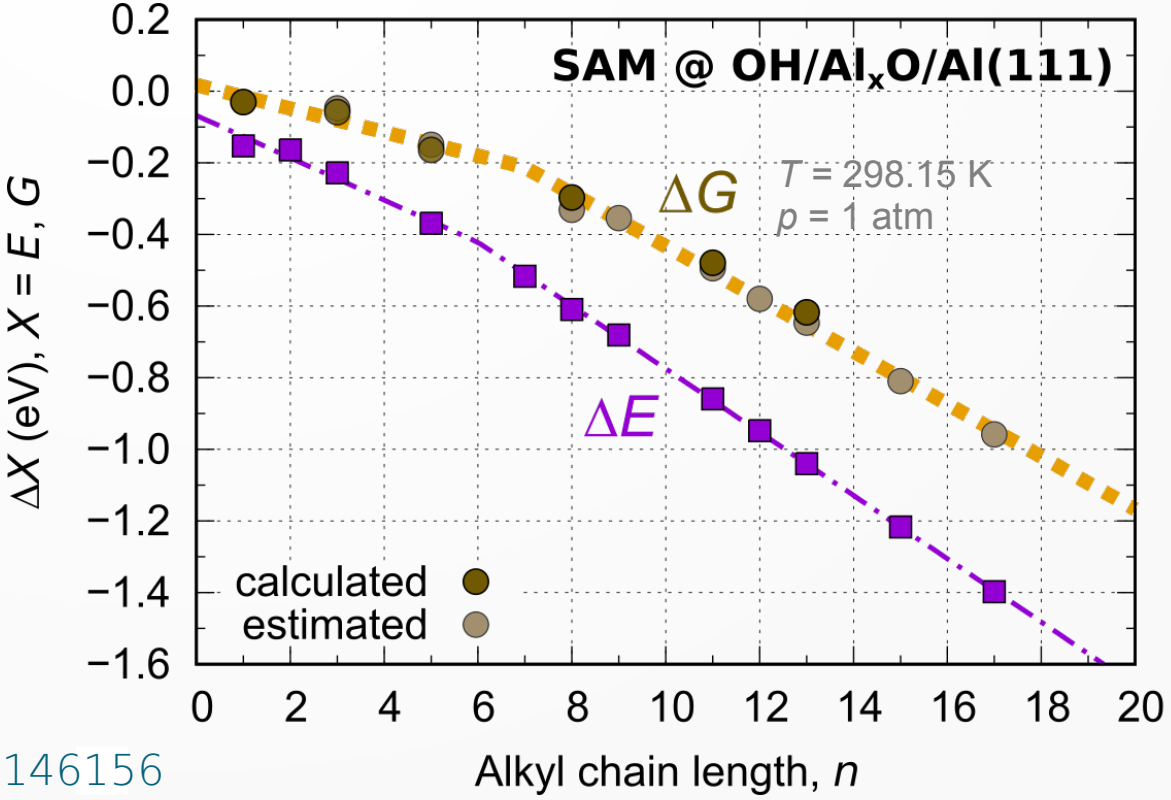


Adsorption energy at full monolayer

1 eV ≈ 100 kJ/mol
≈ 25 kcal/mol

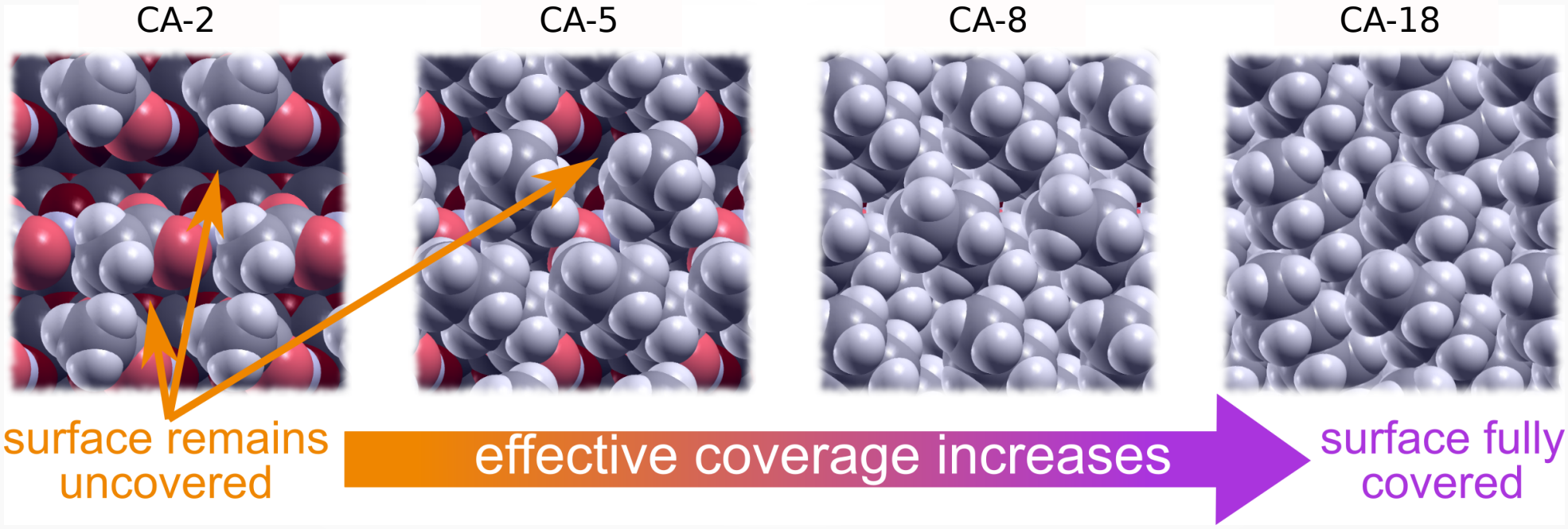


ΔE and ΔG vs. alkyl chain length



Poberžnik et al., Appl. Surf. Sci. **525**, 146156

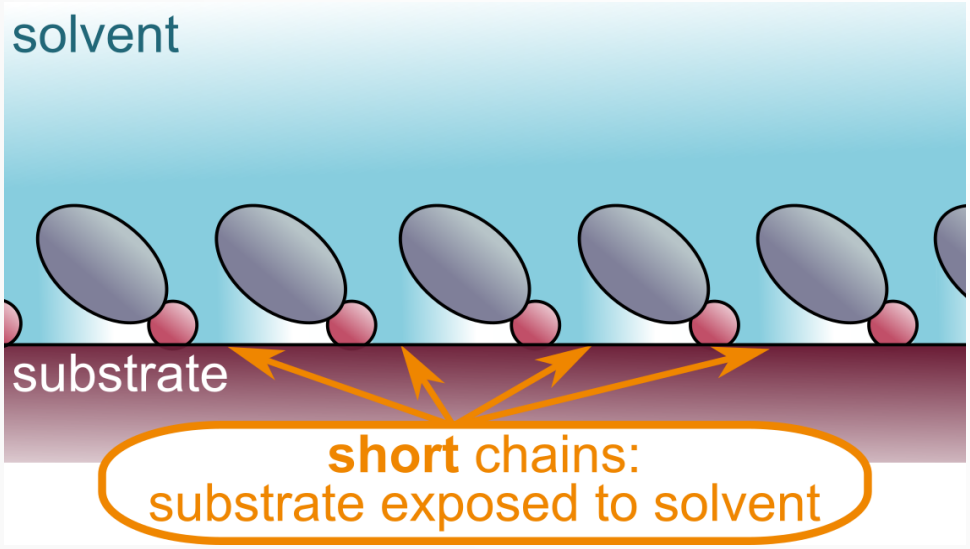
Effective coverage

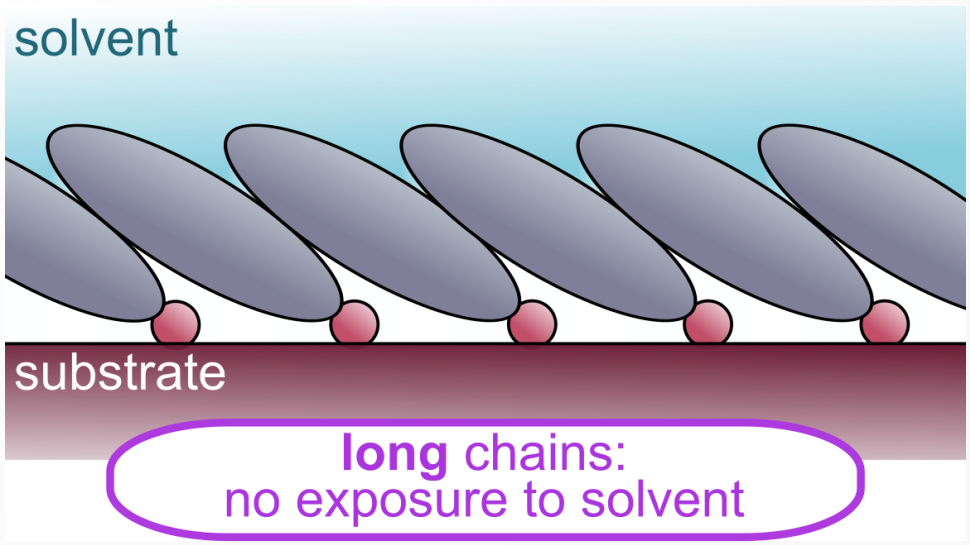


longer alkyl chain = smaller gaps = greater effective coverage

Milošev et al., J. Electrochem. Soc. **166**, C3131

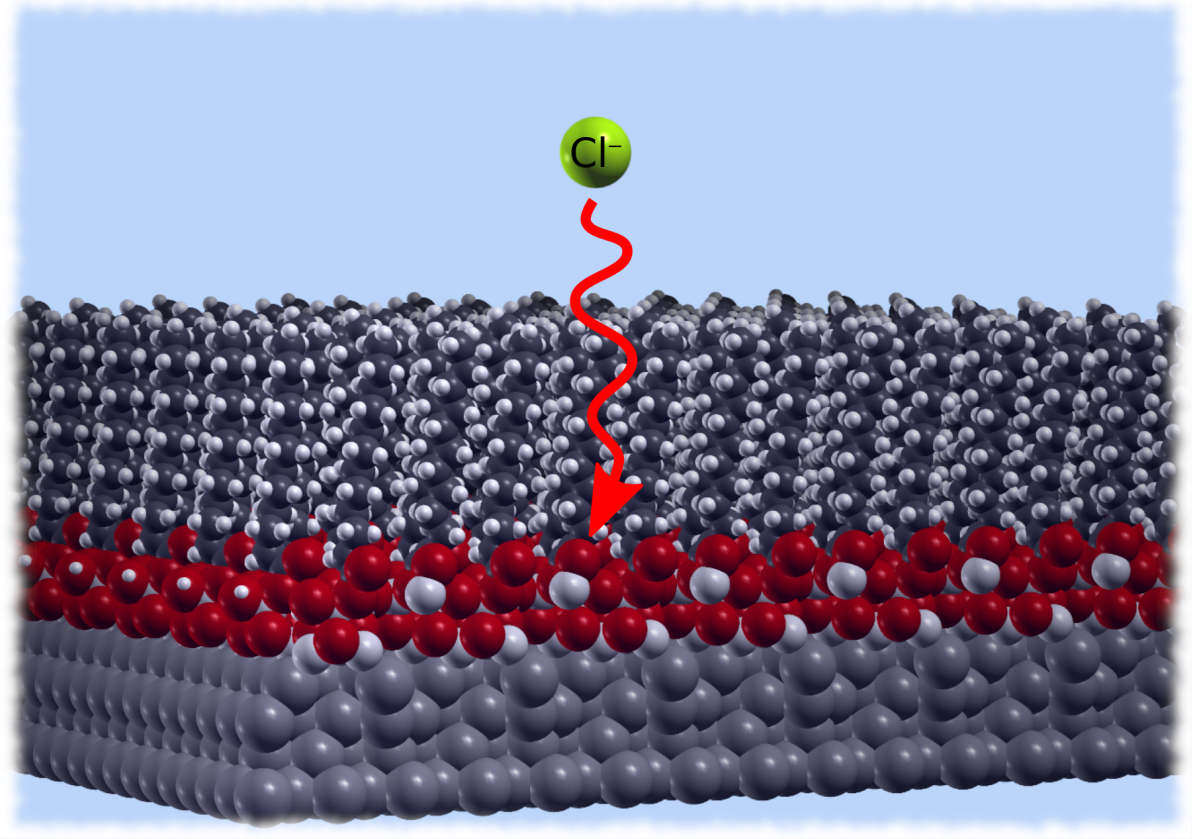
Effective coverage





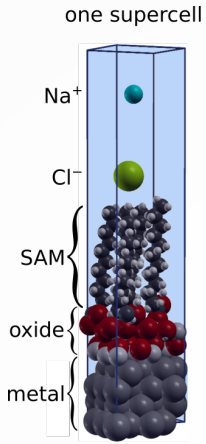
Milošev et al., J. Electrochem. Soc. **166**, C3131

Penetration of Cl^- through SAM



Kokalj & Costa, J. Electrochem. Soc. **168**, 071508

Beware of PBC (periodic-boundary conditions) !



PBC & divergent Coulomb interactions:

add Na^+ to make a supercell neutral

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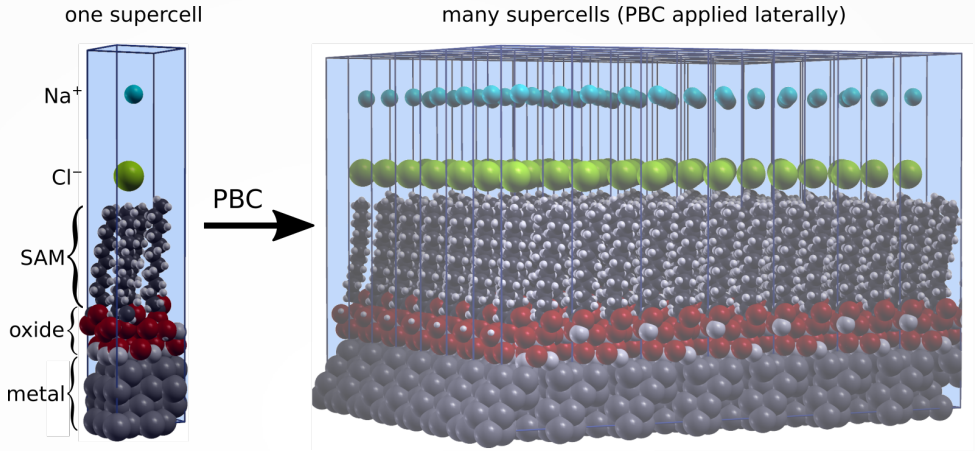

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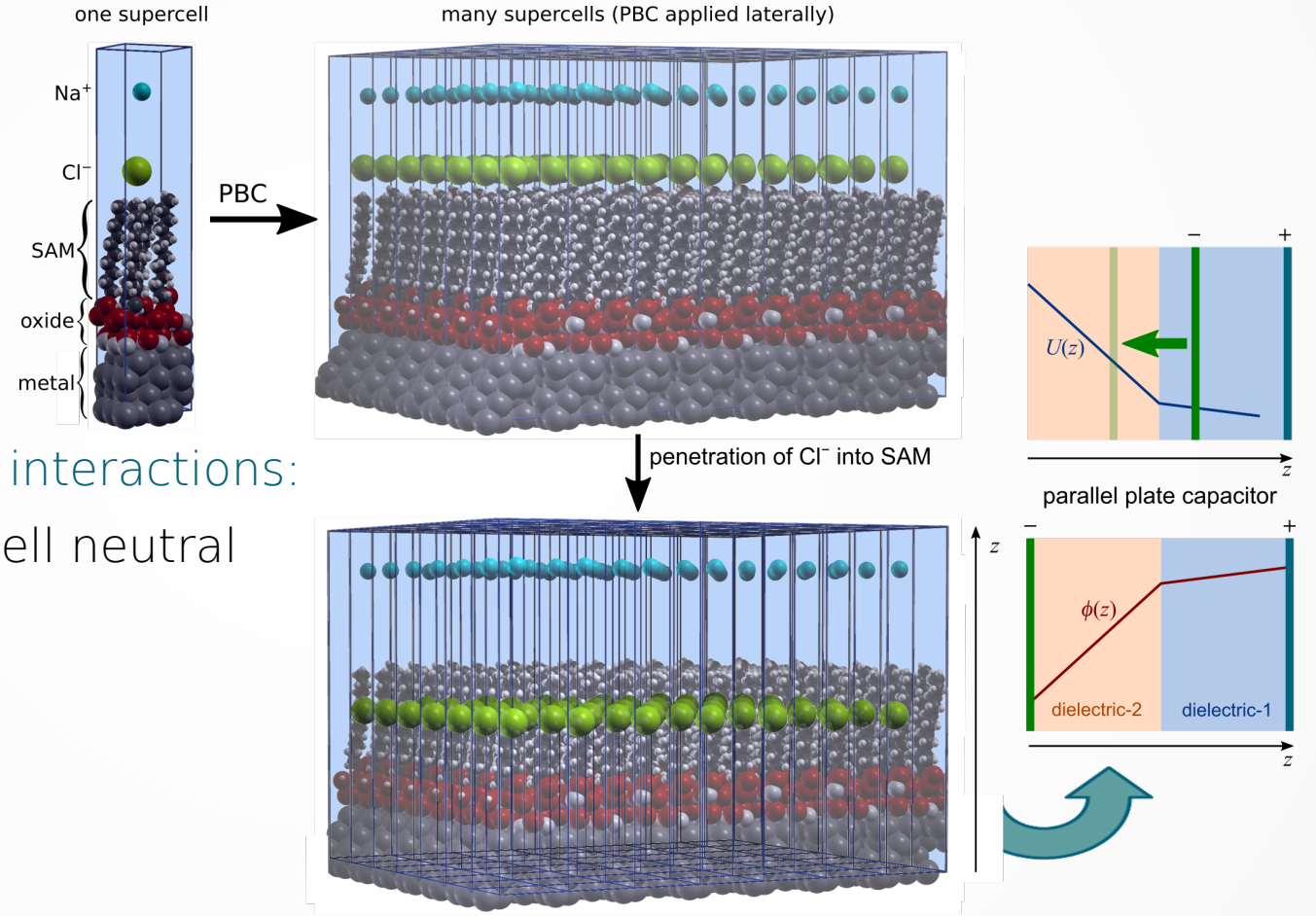

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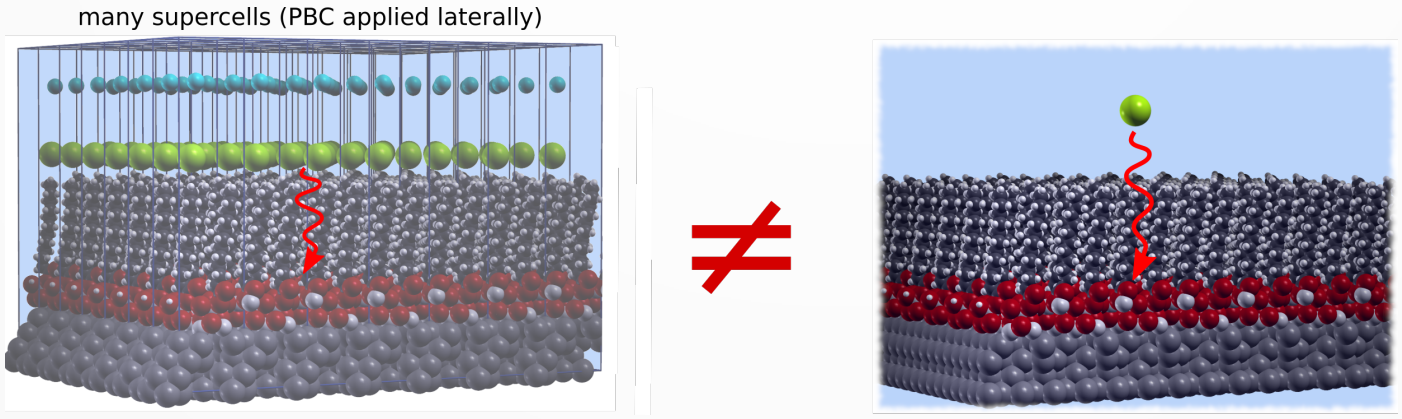
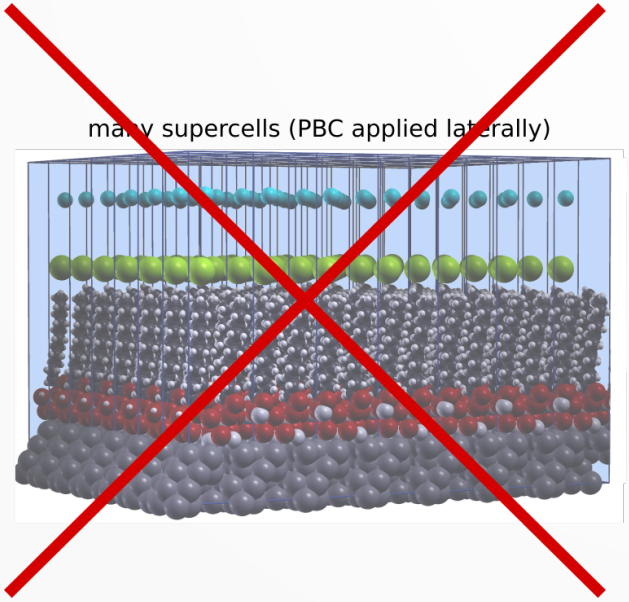
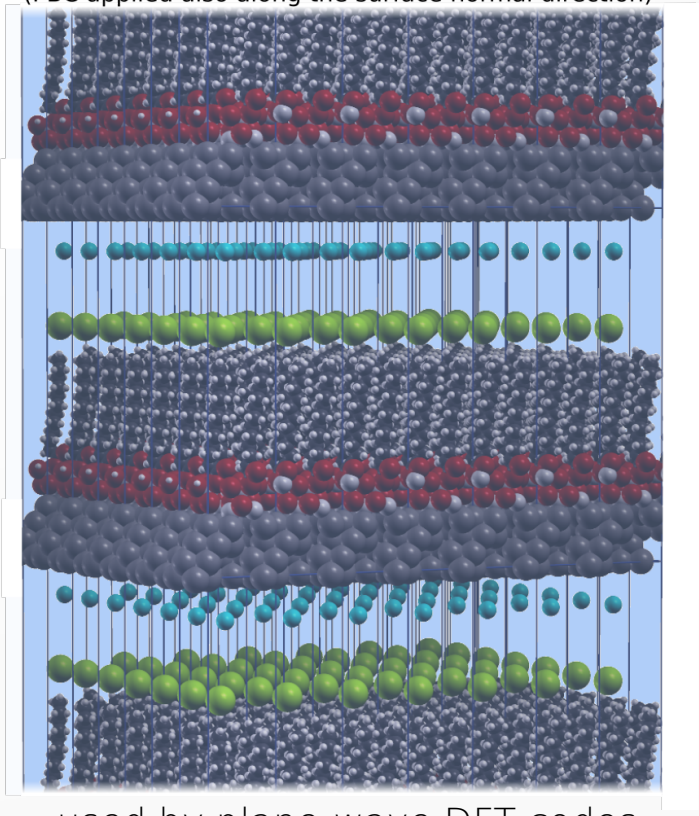


PBC = periodic-boundary conditions

3D PBC: beware of artifacts



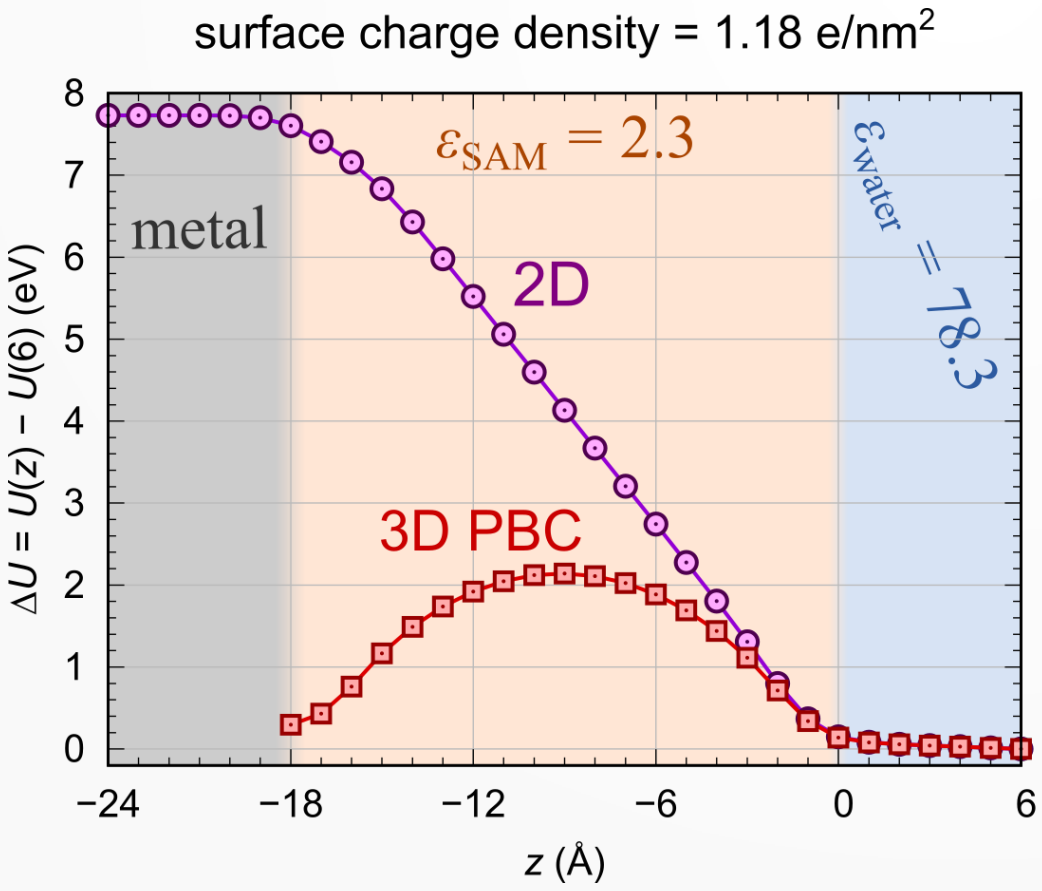
multi-slab model
(PBC applied also along the surface normal direction)



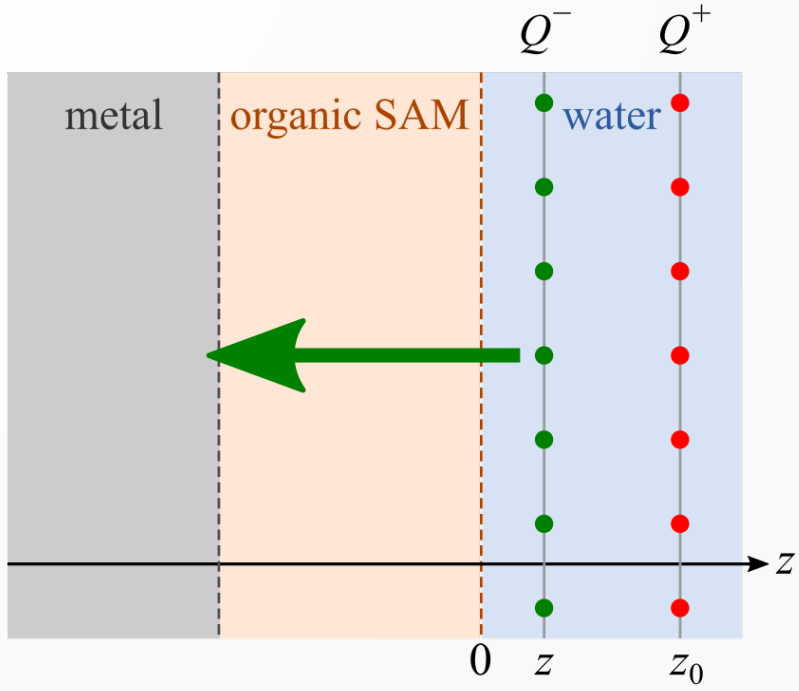
PBC = periodic-boundary conditions

used by plane-wave DFT codes

3D PBC: beware of artifacts



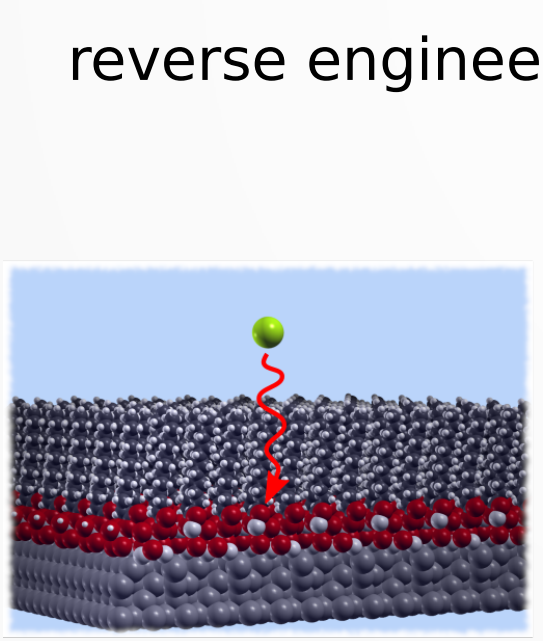
inter-ion electrostatic energy

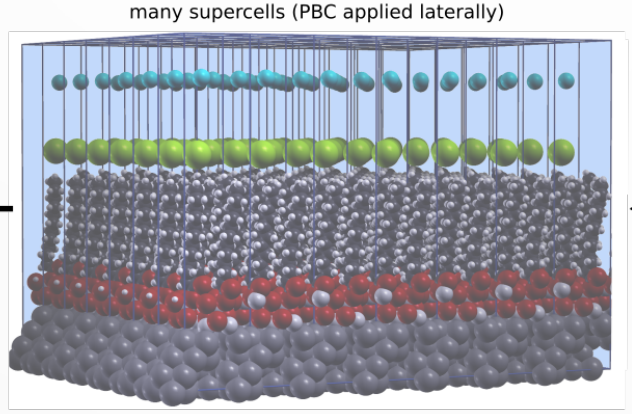


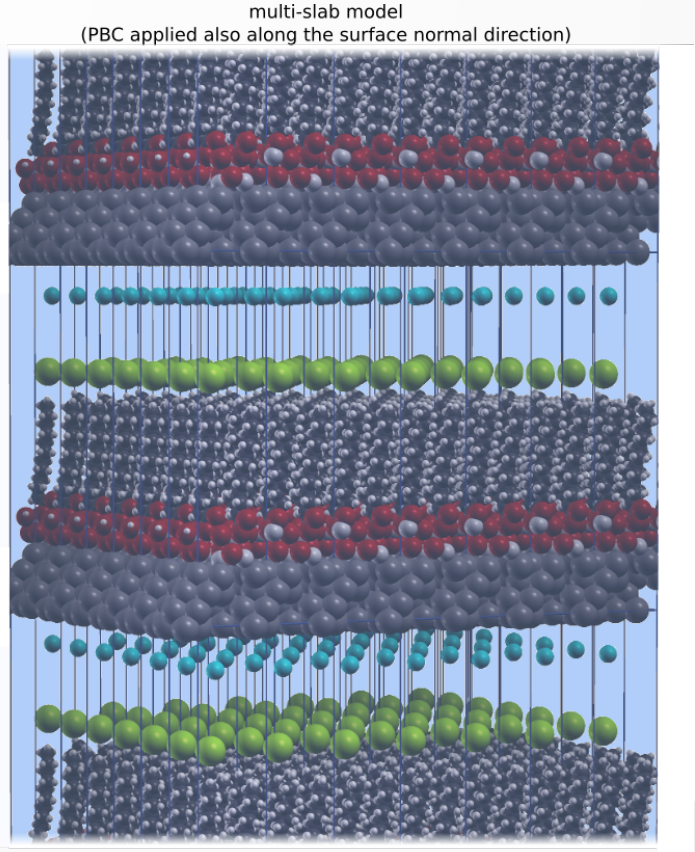
SAM = self-assembled monolayer

3D PBC: beware of artifacts

reverse engineer to get rid of PBC artifacts







PBC = periodic-boundary conditions

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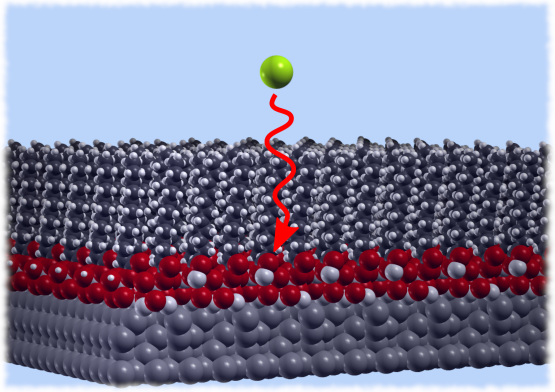
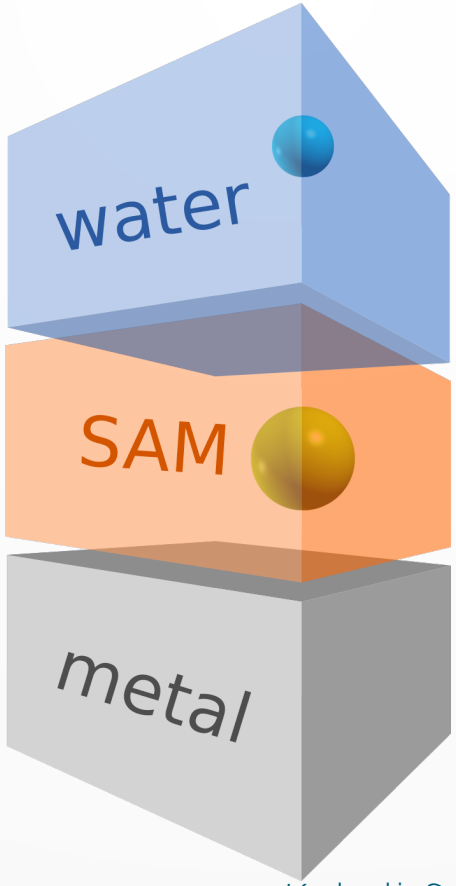
Penetration of Cl^- through SAM





Simplified model calculations:

- metal/SAM/water described by implicit continuum slabs
- ions treated explicitly (ionic solvation considered)
- elastic penalty for Cl^- penetration into SAM neglected

SAM = self-assembled monolayer

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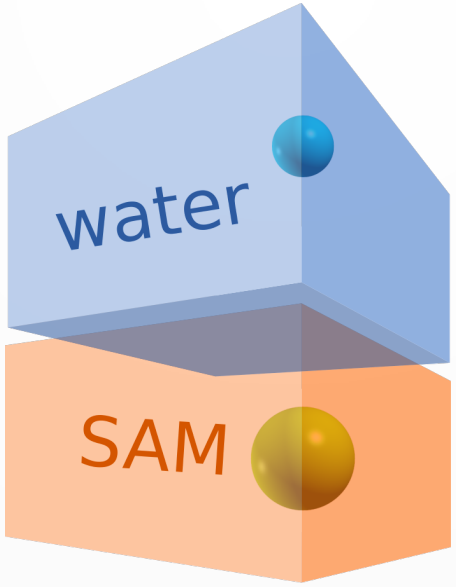
Simpler system: SAM/water



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Environ plugin of Quantum ESPRESSO (QE)
www.quantum-environ.org

SAM = self-assembled monolayer

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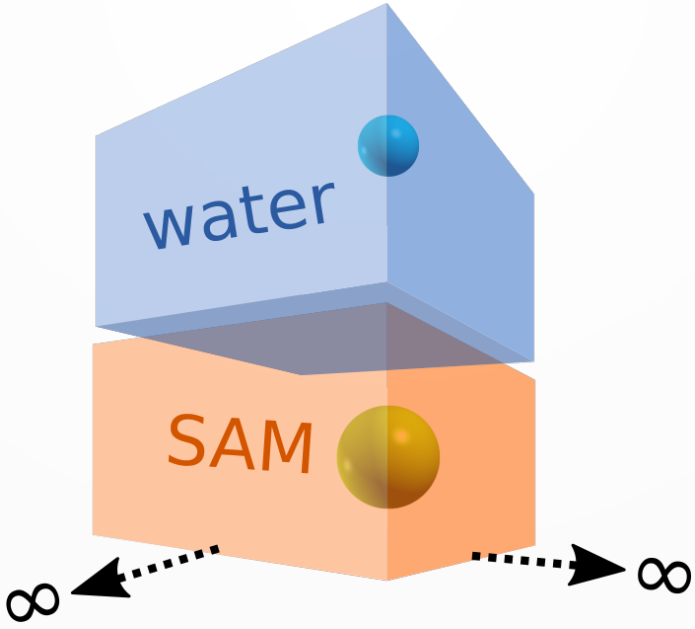

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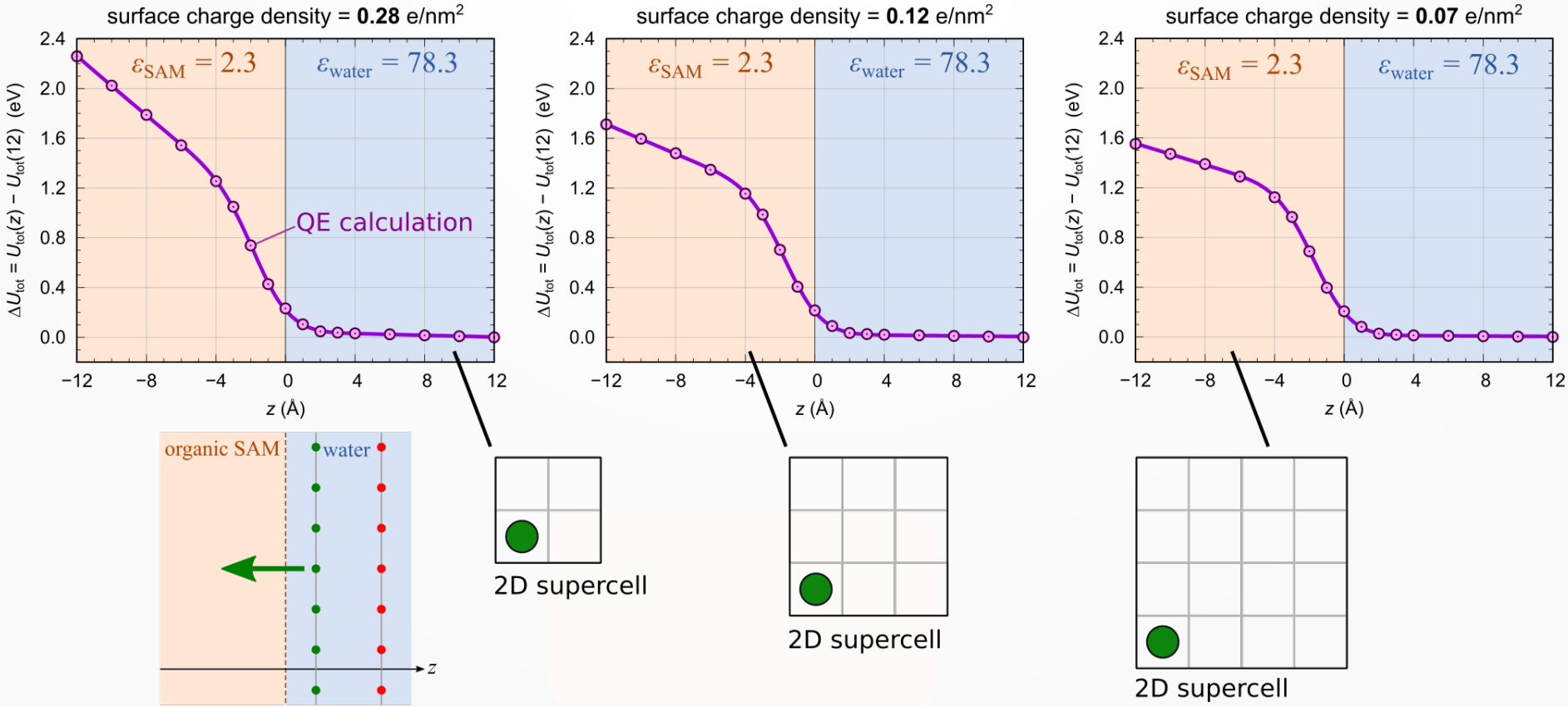

QE = Quantum ESPRESSO

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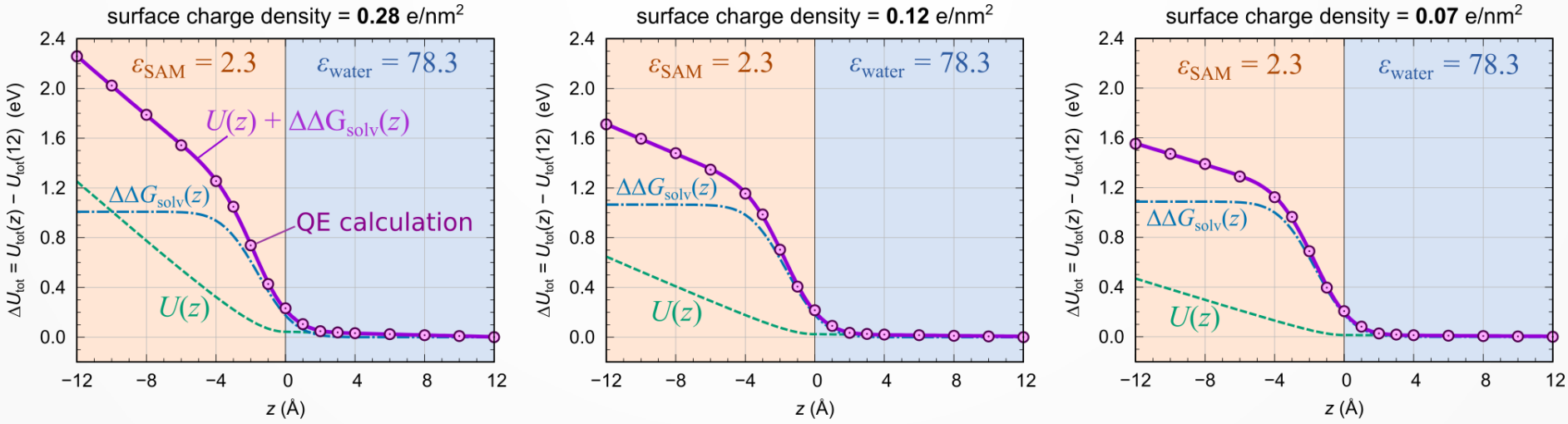
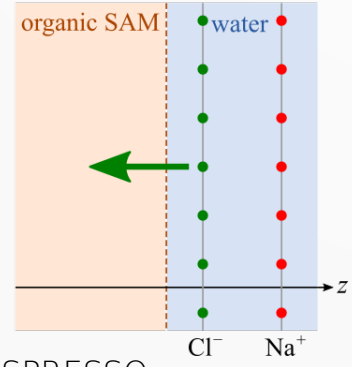
Simpler system: SAM/water





$U(z)$ = electrostatic interactions between ions

$\Delta\Delta G_{solv}(z)$ = difference between ionic solvation (ΔG_{solv}) in SAM and water

QE = Quantum ESPRESSO

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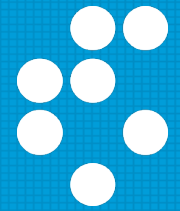
How to understand $\Delta\Delta G_{\mathrm{solv}}$



- Born solvation model:

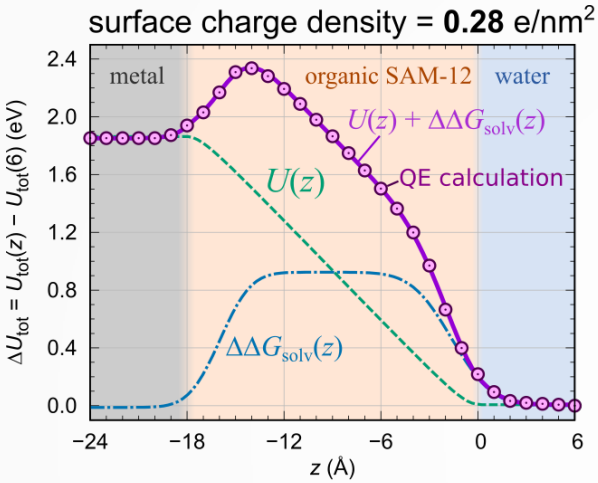
$$\Delta G_{\mathrm{solv}}^{\mathrm{Born}} = -\frac{q_{\mathrm{ion}}^2}{2r_{\mathrm{ion}}}\left(1-\frac{1}{\varepsilon}\right) \quad \text{[in Hartree atomic units]}$$

- SAM vs. water difference:

$$\Delta\Delta G_{\mathrm{solv}}^{\mathrm{Born}} = \Delta G_{\mathrm{solv}}^{\mathrm{Born}}(\mathrm{SAM}) - \Delta G_{\mathrm{solv}}^{\mathrm{Born}}(\mathrm{water}) = \frac{q_{\mathrm{ion}}^2}{2r_{\mathrm{ion}}}\left(\frac{1}{\varepsilon_{\mathrm{SAM}}}-\frac{1}{\varepsilon_{\mathrm{water}}}\right)$$

SAM = self-assembled monolayer

Full system: metal/SAM/water



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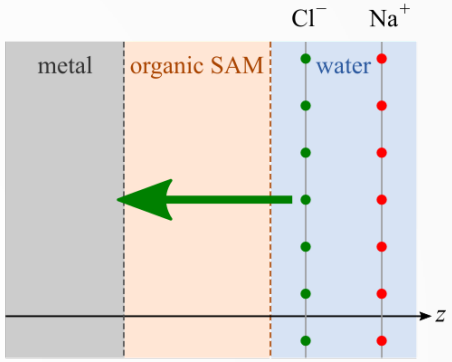

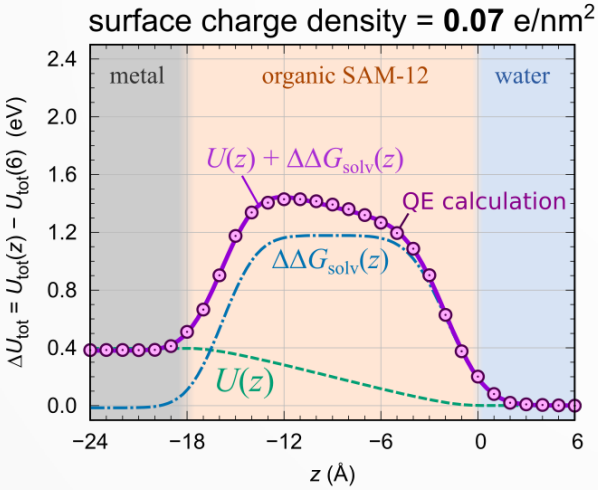

SAM-12 = SAM composed of CA-12
(CA = carboxylic acid)

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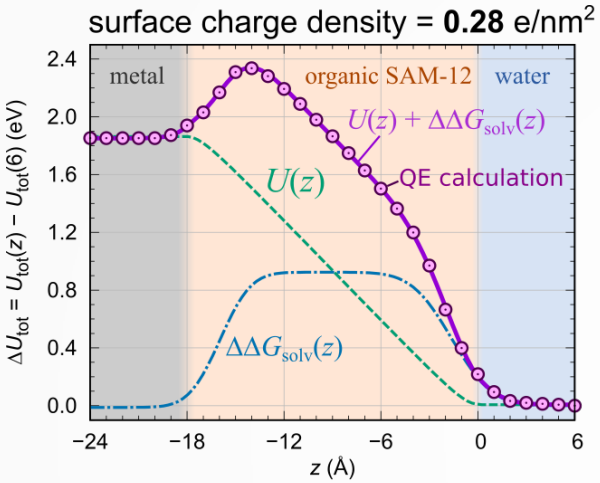
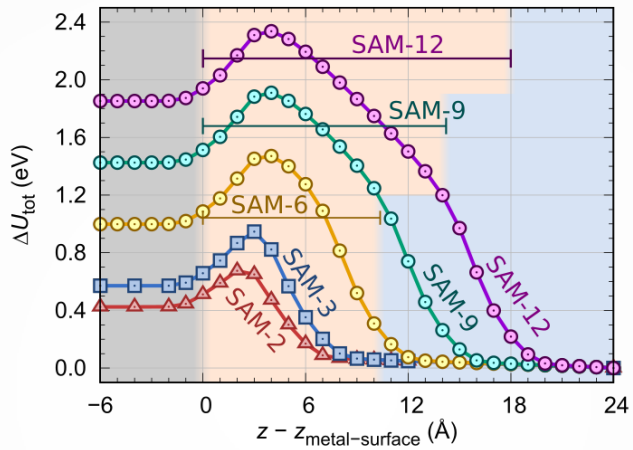
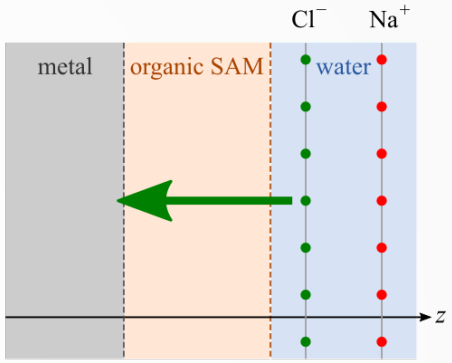
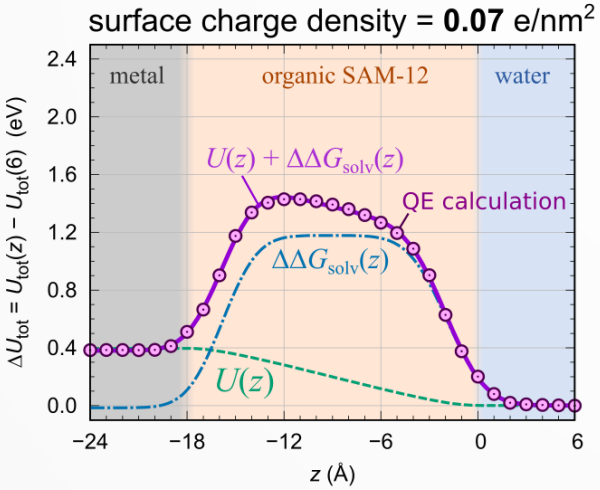
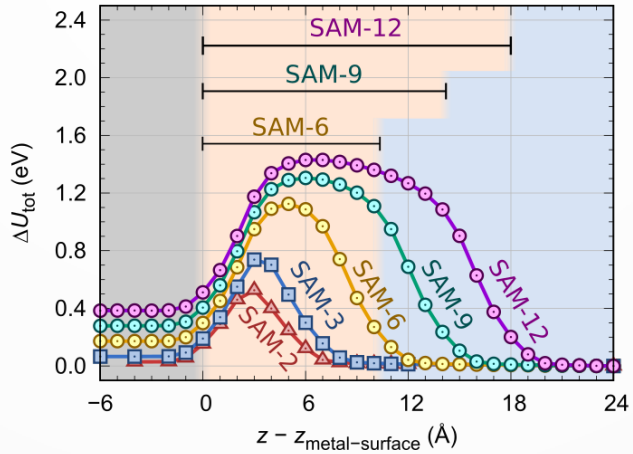
Full system: metal/SAM/water









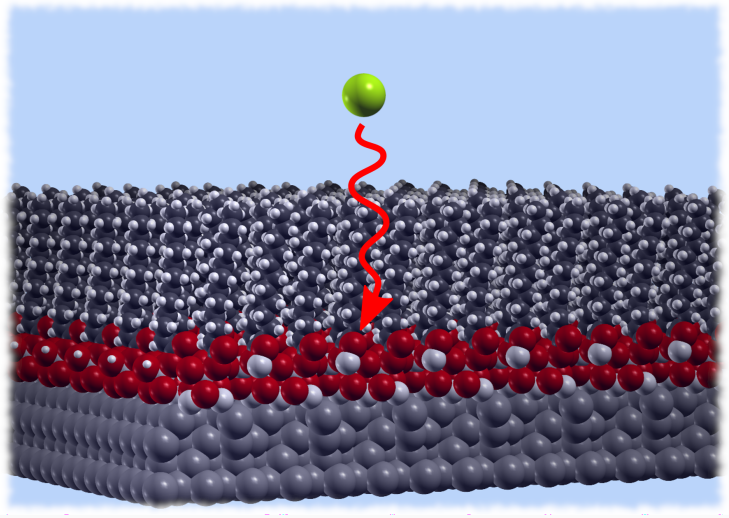


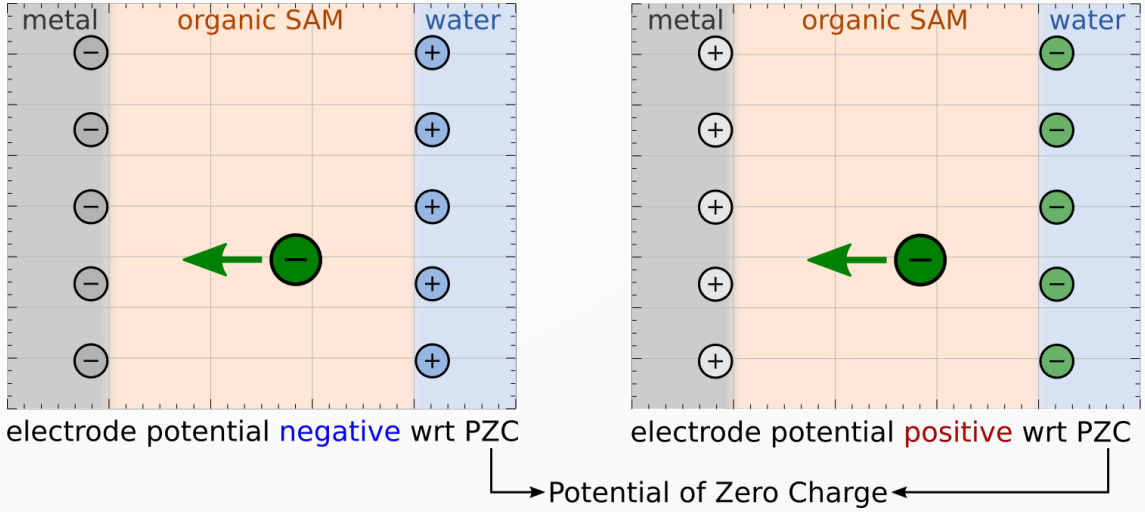
SAM-*x* = SAM composed of CA-*x*
(CA = carboxylic acid)

SAM needs some thickness for the solvation barrier to fully develop !

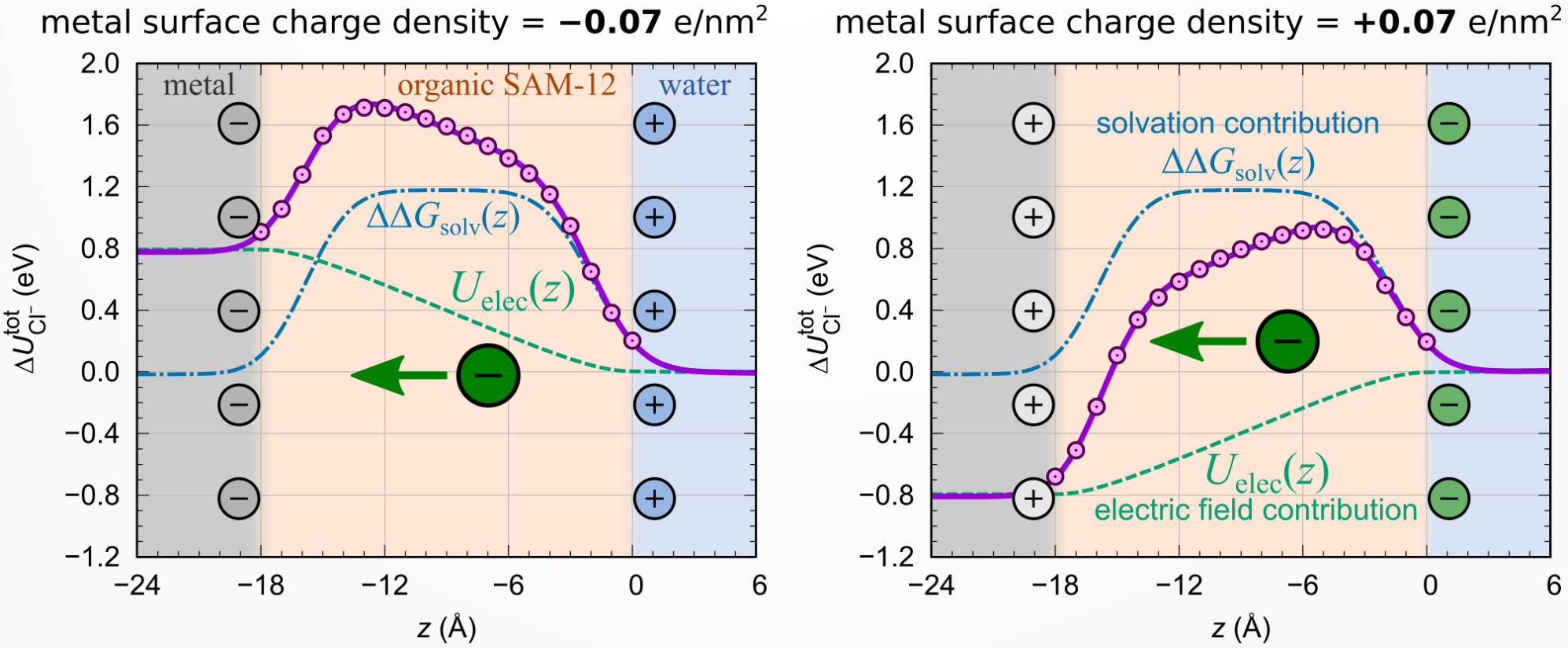
Penetration of a single Cl^-

… as a function of the electrode potential
(a simple Helmholtz double-layer picture)





Penetration of a single Cl^-



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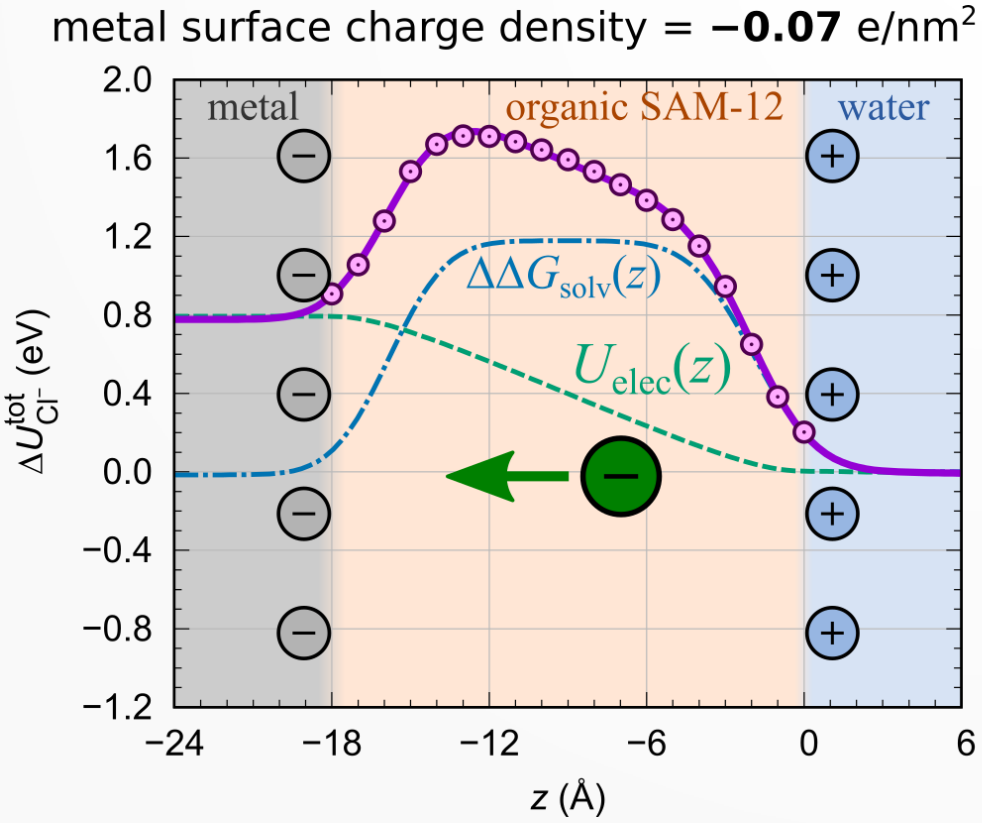
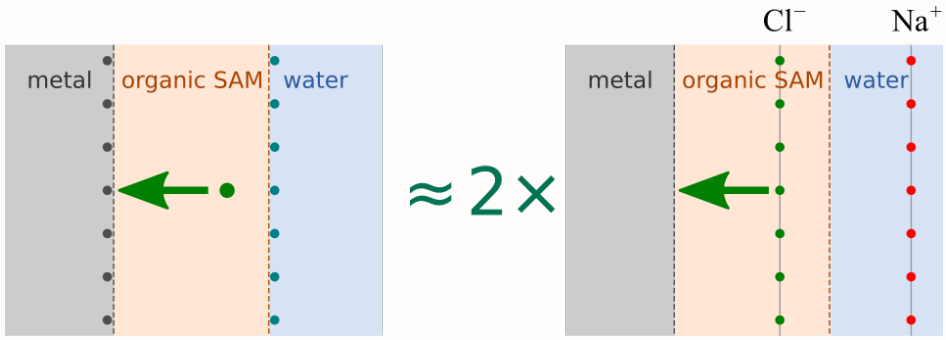
Penetration of a single Cl^-





$$U_{\mathrm{elec}}(z) = 2U(z)$$

elec ≡ electric-field contribution

Note: electrode potential is associated with $U_{\mathrm{elec}}(z)$

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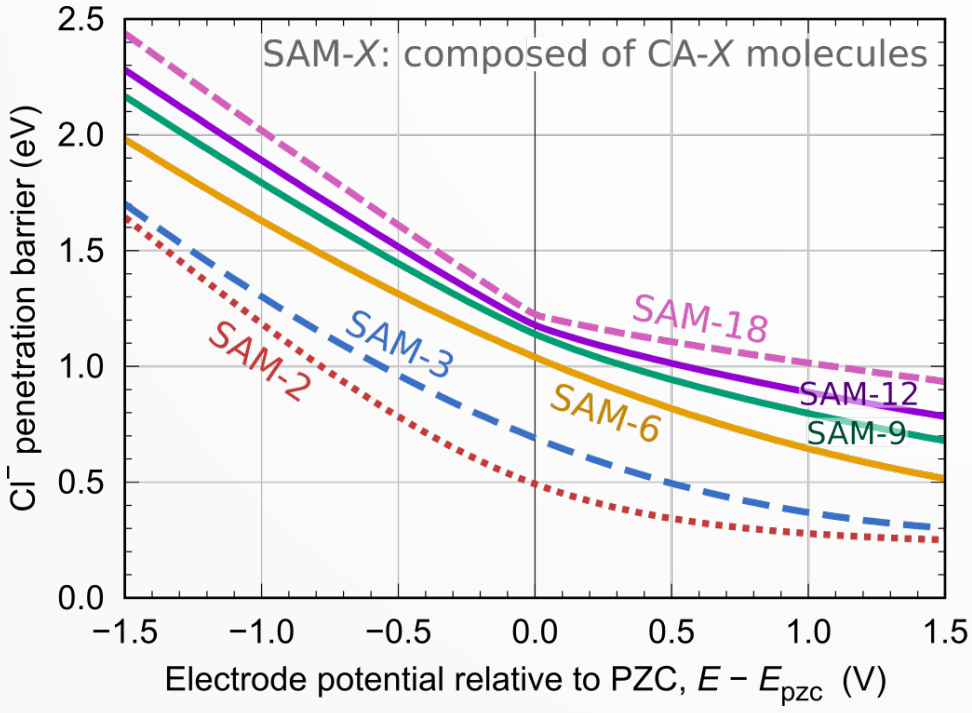
Cl^- penetration vs. electrode potential



Simplifying assumptions:

- CA molecules fully cover the surface irrespective of the alkyl chain length
  (in reality only long-chain CAs form a complete SAM)
- surface is fully covered by SAM at all electrode potentials
- elastic penalty for Cl^- penetration is neglected
- …

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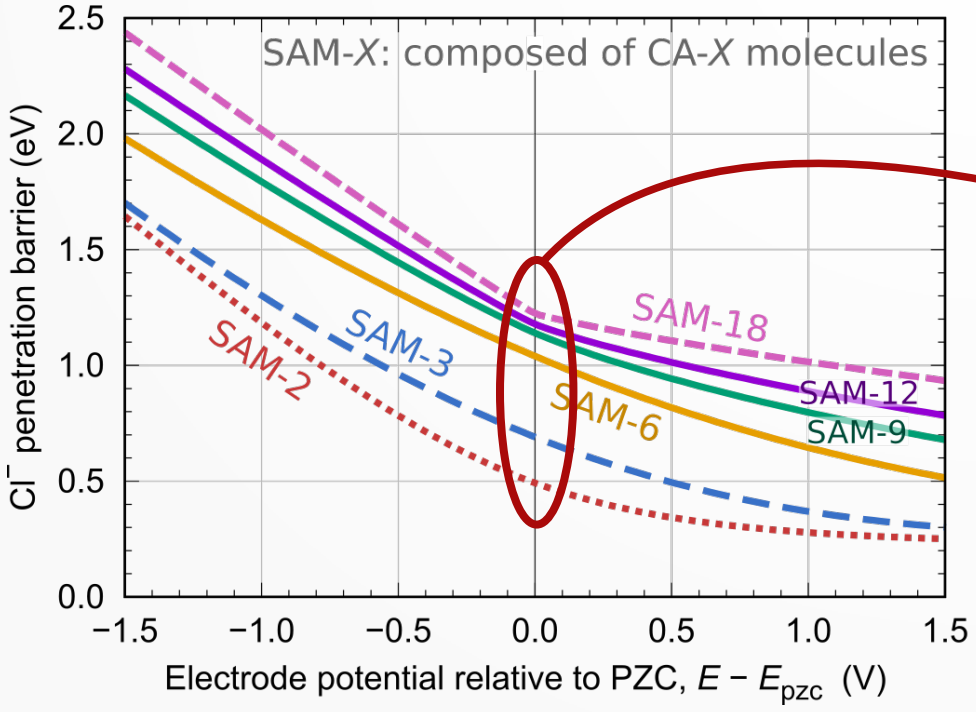
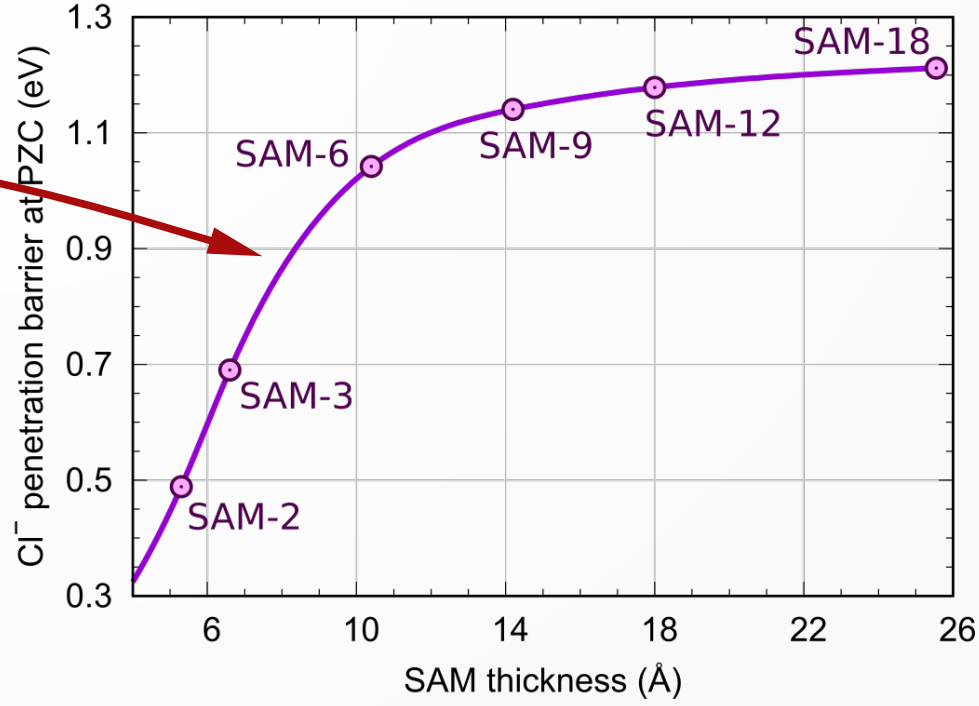
Cl^- penetration vs. electrode potential





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Conclusions

- The mechanism by which SAM hinders the penetration of Cl^- ions into SAM understood with aid of a simple model
- Contributions to penetration barrier:
  - inferior solvation of ions in SAM ($\Delta\Delta G_{solv}$)
  - contribution due to electric field in the Helmholtz double-layer (U_{elec})
  - elastic penalty (currently neglected)
- SAM needs to be thicker than about 10 Å for the barrier to fully develop
- Developed model is based on electrostatic arguments and is **thus generally applicable** (can treat the penetration of any ions, including the dissolution of metals)

Kokalj & Costa, J. Electrochem. Soc. **168**, 071508

Acknowledgments

DFT group:

Matic Poberžnik
Matjaž Dlouhy
Lea Gašparič
Erik Gregori
Matic Lozišek
Dunja Gustinčić
Nataša Kovačević
Sebastijan Peljhan

Electrochemistry group:

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Dževad Kozlica
Ivan Spajić
Dolores Zimerl...

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Slovenian Research Agency

REPUBLIC OF SLOVENIA
MINISTRY OF EDUCATION, SCIENCE AND SPORT

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Thank you for your attention

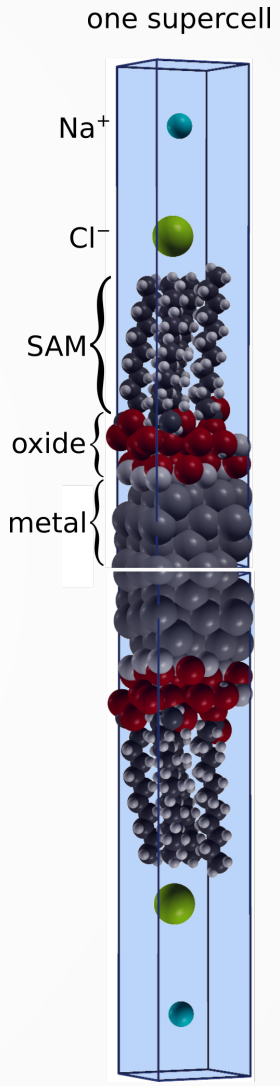
Chimie Paris ParisTech | PSL

Agence Nationale de la Recherche ANR

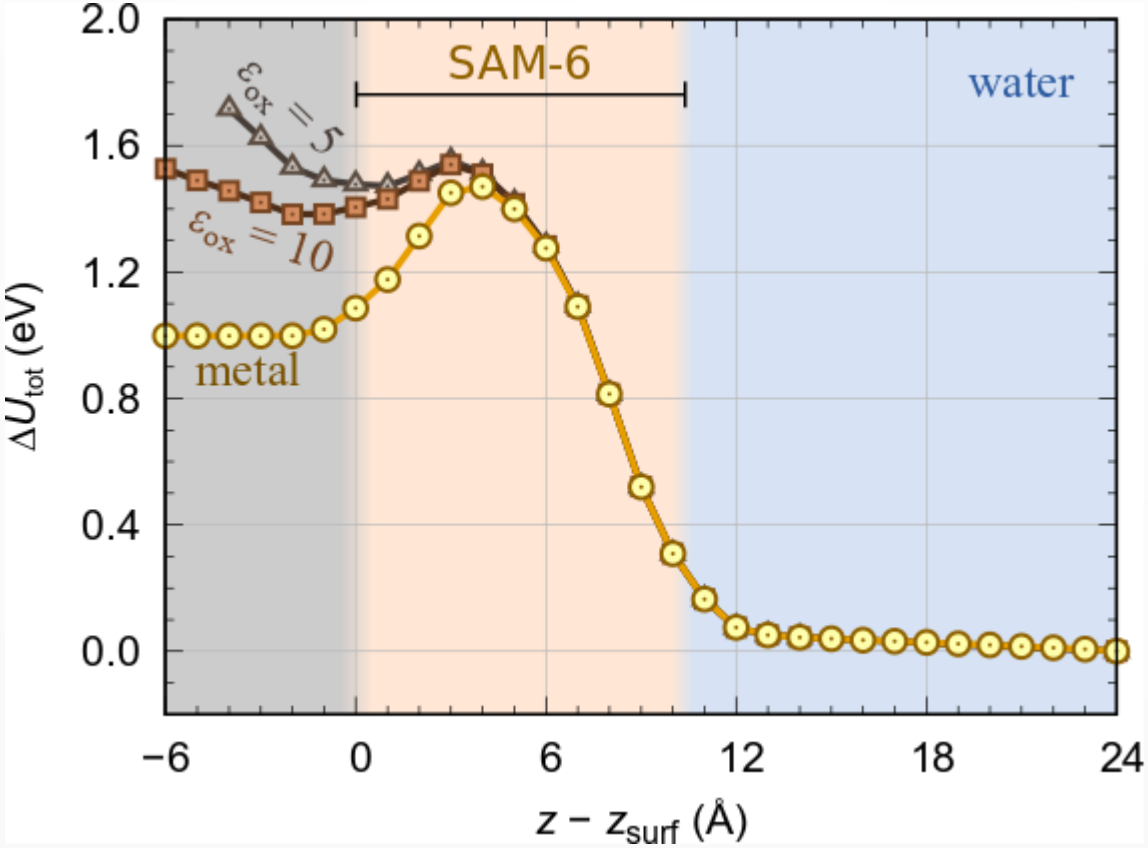
D. Costa
P. Marcus
F. Chiter

How to get rid of 3D PBC artifacts

simple but computationally inefficient method is to use symmetric setup



Oxide/SAM/water



Oxide/SAM/water